



Physical Chemistry-Properties of Gases

70/100
12-01-2021
Abduljabbar
Dr. Abduljabbar Rushdi



Name of a student عبد الجبار رشدي Signature [Signature] No. 6

University of Mustansiriyah

1st Semester-2021

Department of Chemistry

1st Exam-paper A

Q1: Circle the right answer for all of the following:

1: A vessel of 50 mL capacity contains a certain amount of gas at 40 °C and 2 bar pressure. The gas is transferred to another vessel of volume 100 mL at 40 °C. What should be its pressure?

Answer: a) 1.0 atm b) 0.85 mmHg c) 0.9 cmHg d) 1 bar 5/5

2: What is the right formula of the Van der Waals equation?

Answer: a) $p = [nRT/(V-nb)] - n(a^2/V^2)$ b) $P = [nRT/(V-nb)] - V(n^2/a^2)$ c) $p = [nRT/(b-nV)] - a(n^2/V^2)$ d) $P = [nRT/(V-nb)] - a(n^2/V^2)$ 5/5

3: Calculate the temperature of 4.0 mol of a gas occupying 5.0 dm³ at 3.3 bar?

Answer: a) 50.3 °C b) 48 K c) 51 °C d) 50.3 K 5/5

4: Calculate the weight of O₂ (32 g.mol⁻¹) in a 4 L cylinder at 9 atm and 281 K.

Answer: a) 50 kg b) 50 g c) 50 K d) 50 °C 5/5

Q1 40/50

5: Calculate the p_c of He gas, if the p_r and p is 0.44 and 1 atm respectively

Answer: a) 2.26 K b) 2.26 atm c) 2.26 L d) 2.26 mol 5/5

6: If the repulsion forces are negligible, that means the gas is?

Answer: a) real b) noble c) perfect d) compressed 0/5

7: According to the Dalton's law total mole fraction is equal to?

Answer: a) 0.10 mol b) 1.0 mol c) 0.10 d) 1.0 0/5

8: What is the partial pressure of a gas in a mixture if the X_i is 0.5, and the conditions are at STP?

Answer: a) 1.5 Pa b) 0.49 bar c) 0.5 atm d) 0.5 bar 5/5

9: If the value of α is 0.082 then the unit of temperature is?

Answer: a) Kelvin b) Celsius c) Fahrenheit d) no one of these 5/5

10: According to the Avogadro's law the amount of a gas at STP is?

Answer: a) 1.00 mol b) 2.00 mol c) 1.00 L d) 2.00 mol 5/5

Q2: The air inside a flexible 3.5 L container has a pressure of 115 kPa. What should the volume of the container be increased to in order to decrease the pressure to 625 torr?

Q3: A 3 dm³ container holds 0.5 moles of N₂ gas at 42 °C. What is the pressure inside the container?

Q2

- لک

$$V_1 = 3.5 \text{ L}, P_1 = 115 \text{ kPa}$$

$$P_2 = 625 \text{ torr}, V_2 = ?$$

تحوّل إلى atm

$$P_1 = \frac{115 \text{ kPa}}{101.25 \text{ atm}} = 0.011 \text{ atm}$$

$$P_2 = \frac{625 \text{ torr}}{760} = 0.822 \text{ atm}$$

$$P_1 V_1 = P_2 V_2$$

$$Q_2 \frac{15}{25}$$

$$0.011 \text{ atm} \times 3.5 \text{ L} = 0.822 \text{ atm} \times V_2$$

$$0.0385 \text{ atm} \cdot \text{L} = 0.822 \text{ atm} \times V_2$$

$$V_2 = \frac{0.0385 \text{ atm} \cdot \text{L}}{0.822 \text{ atm}}$$

$$V_2 = 0.04 \text{ L}$$

This the reason why

Your Result is not ~~right~~ right

take care next time!

Q3

لک

$$n = 0.5 \text{ mol}, V = 3 \text{ L}$$

$$T = 42^\circ \text{C}, P = ?$$

$$R = 0.082 ?$$

$$315 \text{ K} = 42 + 273 = T$$

$$PV = nRT$$

$$\frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}$$

$$P(3 \text{ L}) = 0.5 \text{ (mol)} \times 0.082 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \times 315 \text{ K}$$

$$P(3 \text{ L}) = 12.915 ?$$

$$P = \frac{12.915}{3}$$

$$P = 4.30 \text{ atm}$$

$$Q_3 \frac{15}{25}$$