50/100 Fifty only

12-01-2021

Physical Chemistry-Properties of Gases

Dr. Abdulljabbar R. Rushdi

Name of a student ---------- Signature ---------- No. 13

University of Mustansiriyah — **1st Semester-2021**

Department of Chemistry — **1st Exam-paper B**

Q1: Circle the right answer for all of the following:

1: Helium represents a.

Answer: a) real gas b) ideal gas c) noble gas d) heavy gas

2: A 0.2 L container contains a certain amount of gas at 1.0 bar pressure. The gas is transferred to another vessel of volume 0.5 dm^3. What should be it is pressure?

Answer: a) 0.60 atm b) 0.40 dm^3 c) 0.4 atm d) 0.4 mmHg

3: A gas occupies 299 dm^3 at 127 °C and 760 mm pressure. What would be it is volume at STP?

Answer: a) 199.8 L b) 199 dm^3 c) 200 L d) 204 dm^3

4: Calculate the weight of CH_4 (16 $g.mol^{-1}$) in a 10 L cylinder at 15 atm and 34 °C.

Answer: a) 95.33 g mol^{-1} b) 95.33 g c) 85.80 mol d) 86.65 g

5: Calculate the number of moles for CH_4 in a 12 L cylinder at 14 bar and 28 °C.

Answer: a) 6.8 mol b) 6.9 mol c) 6.5 mol d) 6.7 mol

6: According to Graham's law the heaviest gas is?

Answer: a) H_2 b) O_2 c) N_2 d) CO_2

7: According to the Avogadro's law the amount of a substance is directly proportional with?

Answer: a) p b) T c) R d) V

NO ANSWAR

8: The difference between real and ideal gas is one of the following?

Answer: a) p & V b) T & n c) d) attraction forces & volume of a gas

9: It can know the molecular mass of un known gas by applying one of the following?

Answer: a) Boyle's law b) Graham's law c) Charles's law d) Gay-Lussac's law

10: If V_m is bigger than V^o_m then this means the behaviour of a gas is?

Answer: a) Real b) Ideal c) Real & ideal d) Z = 0

Q2: A gas sample has a mass of 9.98 g. Its volume is 21.6 L at a temperature of 75.46 °C and a pressure of 641 Torr. Calculate its molar mass.

Q3: A 1.3 mole of Ar gas is placed in a container at 27 °C at a pressure of 725 torr. What is the volume of the container in ml?

12/01/2021 — **Best wishes** — **Dr Abduljabbar I. R. Rushdi**

Q2 wt = 0.98g V = 21.6L , T = 75.46C P = 641 torr / 1atm

mwt = ?

75.46 + 273

$PV = \frac{wt}{mwt} RT \Rightarrow 641\ atm * 21.6L = \frac{0.98g}{mwt} * 0.082 \ldots$

$641\ atm * 21.6L = \frac{0.98g}{mwt} * 0.082\ atm*L/mol*K \times 348.46\ K$

$mwt = \frac{0.98g * 0.082\ atm*L/mol*K * 348.46K}{641\ atm * 21.6L}$

mwt = 0.9436 g/mol

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Q3// n = 1.3mol T = 27C° P = 725 toor / 1atm V = ?

PV = nRT

T = 27 + 273 = 298 K (301, 303K)

725 atm * V = 1.3mol * 0.082 atm*L/mol/K * 298 K

$V = \frac{1.3mol * 0.082 \frac{atm*L}{mol*K} * 298K}{725 atm}$

Q3 5/25

V = 0.043 L ⇒ 0.043 X 1000 = 43.8mL