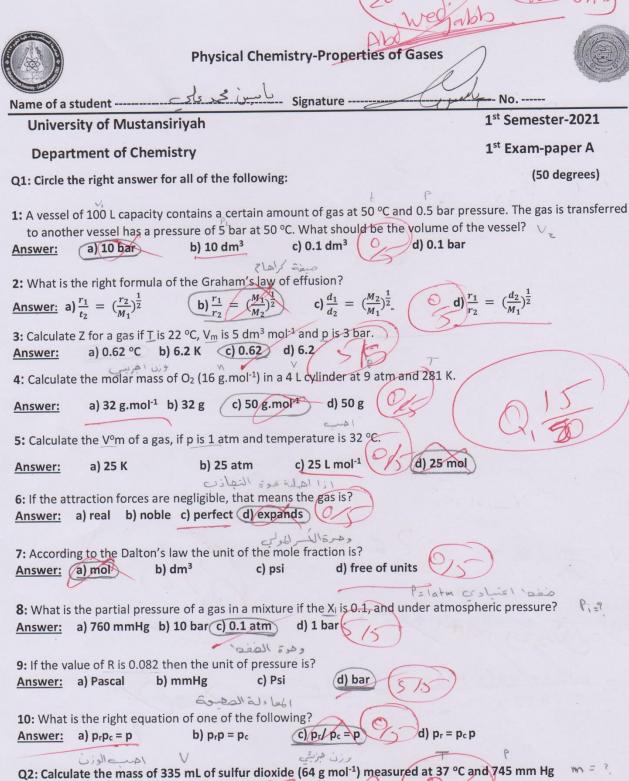
1st Semester-2021 1st Exam-paper A (50 degrees) d) 25 mol d) free of units



Wed 20/01/2021

(12 2 24) ils 200)

Best wishes

Q3: Calculate the volume of 0.25 g of oxygen at 25 °C and 742 mm Hg pressure.

Dr Abduljabbar I. R. Rushdi

(25 degrees)

(25 degrees)

$$PV = \frac{m}{M} RT$$
= units

6.9 × 335 = 6.75 × 0.08 2 × (25+273)

$$M = \frac{0.25 \times 0.082 \times 2.58}{0.7 \times 3.75} = \frac{6.(0.9)}{3.01.5} = \frac{6.(0.9)}{3.01.5} = \frac{6.(0.9)}{3.01.5}$$