Name of a student ------محمد زهير وليد مكبوح------ Signature ---mehembe--- No. ---30---

University of Mustansiriyah   **1st Semester-2021**

Department of Chemistry   **1st Exam-paper B**

Q1: Circle the right answer for all of the following: (50 degree)

1: Carbon dioxide is classified as a .

Answer: a) toxic gas   b) ideal gas   (c) real gas   d) heavy gas

2: A 2 dm^3 container contains a certain amount of gas at 0.5 atm pressure. The gas is transferred to another vessel of volume and the pressure is 0.25 bar. What should be it is Volume?

Answer: a) 0.40 atm   b) 0.40 dm^3 (c) 0.4 bar   d) 4 bar

3: A gas occupies 400 dm^3 at 130 °C and 76 cmHg pressure. What would be it is volume at STP?

Answer: (a) 270 L   b) 207 dm^3 c) 207 m^3 d) 204 cm^3

4: Calculate the weight of H_2 (2.00 g.mol^{-1}) in a 2 L cylinder at 2.5 atm and 27 °C.

Answer: a) 0.40 mol^{-1} (b) 0.40 g   c) 0.40 mol g^{-1} d) 0.4 g mol^{-1}

5: Calculate the number of moles for CO_2 in a 10 L cylinder at 8 bar and 27 °C.

Answer: a) 3.25 mmol   b) 3.00 mol   c) 3.00 L   d) 2.99 mol

6: According to Graham's law the lightest gas is?

Answer: a) H_2 b) O_2 c) N_2 d) CO_2

7: According to the Boyle's law the pressure of a gas is inversely proportional with?

Answer: a) mol   b) T   c) R   (d) V

8: If a gas has $V_m \neq V^o_m$ then this means one of the following?

Answer: a) real   b) noble   c) ideal   d) heavy

9: If $\underline{RT} > pV$ this means the forces dominated are?

Answer: a) attraction   b) repulsion   (c) Van der Waal's   d) no one of these

10: According to Gay-Lussac's law the volume of the gas is?

Answer: a) constant   b) variable   c) equal to zero   (d) equal to 22.4 L

Q2: Under the same conditions of temperature and pressure, how many times faster will hydrogen effuse

compare to carbon dioxide.   (25 degree)

Q3: Calculate the density of carbon dioxide (44 g mol^{-1}) at STP.   (25 degree)

Wed_20/01/2021   Best wishes   Dr Abduljabbar I. R. Rushdi

Q1/

$r = PVT$

$r = 1 * 1 * 100 k$

$r = 100$ %

Q2/

$$d = \frac{PM}{RT}$$

$$d = \frac{1\,atm * 44\,g/mol^{-1}}{0.082\,L/atm\,mol/K * 273\,K}$$

$$d = \frac{44\,g}{22.386\,L} \implies \boxed{d = 1.9655\,g/L}$$