



Physical Chemistry\_Chpt\_One\_Properties of Gases

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1<sup>st</sup> Exam-paper A

Q1: Circle the right answer for all of the following:

(50 points)

1: If a gas has a non-polar particle then the difference between the volume of this gas is:

Answer: a)  $V_{Real} > V_{Perfect}$  b)  $V_{Real} < V_{Perfect}$  c)  $V_{Real} = V_{Perfect}$  d)  $V_{Real} \neq V_{Perfect}$

2: A gas occupies 300000 mL at 130 °C and 760 mmHg pressure. What would be its volume at STP?

Answer: a) 203.22 mL b) 203.22 dm<sup>3</sup> c) 204 L d) 204 dm<sup>3</sup>

3: Calculate the weight of CH<sub>4</sub> (16 g.mol<sup>-1</sup>) in a 10 L cylinder at 15 x 10<sup>5</sup> Pa and 307 K.

Answer: a) 95.33 g mol<sup>-1</sup> b) 95.33 g c) 95.33 mol d) 95.33 kg

4: Calculate the number of moles for CH<sub>4</sub> in a 10000 mL cylinder at 10<sup>6</sup> Pa and -32 °C.

Answer: a) 4.5 mol b) 4.0 mol c) 4.0 mmol d) 4.5 mmol

5: According to Graham's law the heaviest gas is?

Answer: a) H<sub>2</sub>O b) CH<sub>4</sub> c) NH<sub>3</sub> d) CO

6: A 20 L tank contains a certain amount of gas at 10<sup>5</sup> Pa. The gas is transferred to another tank 40 dm<sup>3</sup>. What should be its pressure?

Answer: a) 0.50 atm b) 50 dm<sup>3</sup> c) 50 atm d) 0.50 mmHg

7: According to the Avogadro's law the amount of a substance is directly proportional with?

Answer: a) p b) T c) R d) V e) n

8: The difference between real and ideal gas is one of the following?

Answer: a) law p & high T b) high p & law T c) high p & high T d) law p & law T

9: It can know the density of a gas by applying one of the following?

Answer: a) Van der Waal's law b) Graham's law c) Charles's law d) Gay-Lussac's law

10: If  $V_m$  is bigger than  $V_m^0$  then this means the behaviour of a gas is?

Answer: a) Real b) Ideal c) Real & ideal d)  $Z < 1$

Q2: A (28 mol) gas sample has a mass of 10000 mg. The volume of a container is 22 dm<sup>3</sup> at a temperature of 76

°C and a pressure of 641 Torr. What is the density of the gas?

(25 points)

Q3: An Ar gas is placed in a container at 30 °C at a pressure of 730 torr. What is the volume of the container in ml?

(25 points)

Q2/

$$PV = nRT$$

$$P_m = \frac{m}{V} RT$$

$$641 = \frac{m}{V} \times 0.082 \times 349$$

$$641 = \frac{10}{22} \times 0.082 \times 349$$

$$P_m = 1.99 \times 10^{-3} \text{ dm}^3$$

$$n = 28 \text{ mol}$$

$$M = \frac{10000}{1000} = 10 \text{ g}$$

$$V = 22 \text{ dm}^3$$

$$T = 76 + 273 = 349 \text{ K}$$

$$P = 641 \text{ Torr}$$

Q2  $\frac{5}{25}$

Q3  $\frac{0}{25}$