

Republic of Iraq  
Ministry of Higher Education and Scientific Research  
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Collage of Science  
Department of Chemistry



# *Practice Qualitative Chemical Analysis*

*First Grade - First Term*

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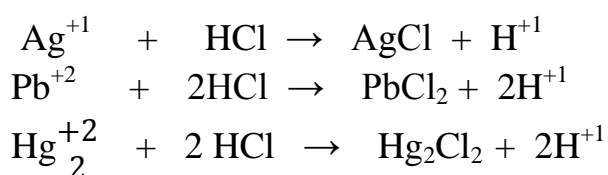
THE PERIODIC TABLE OF THE ELEMENTS																	
1 1A H Hydrogen 1.008	2 2A He Helium 4.003	3 Li Lithium 6.941	4 Be Beryllium 9.012	5 B Boron 10.811	6 C Carbon 12.011	7 N Nitrogen 14.007	8 O Oxygen 15.999	9 F Fluorine 18.998	10 Ne Neon 20.180	11 Na Sodium 22.990	12 Mg Magnesium 24.305	13 Al Aluminum 26.987	14 Si Silicon 28.086	15 P Phosphorus 30.974	16 S Sulfur 32.066	17 Cl Chlorine 35.453	18 Ar Argon 39.948
19 K Potassium 39.098	20 Ca Calcium 40.078	21 Sc Scandium 44.956	22 Ti Titanium 47.88	23 V Vanadium 50.942	24 Cr Chromium 51.996	25 Mn Manganese 54.938	26 Fe Iron 55.845	27 Co Cobalt 58.933	28 Ni Nickel 58.693	29 Cu Copper 63.546	30 Zn Zinc 65.39	31 Ga Gallium 69.723	32 Ge Germanium 72.64	33 As Arsenic 74.922	34 Se Selenium 78.972	35 Br Bromine 79.904	36 Kr Krypton 83.8
37 Rb Rubidium 85.468	38 Sr Strontium 87.62	39 Y Yttrium 88.906	40 Zr Zirconium 91.224	41 Nb Niobium 92.906	42 Mo Molybdenum 95.95	43 Tc Technetium 98.907	44 Ru Ruthenium 101.07	45 Rh Rhodium 102.906	46 Pd Palladium 106.42	47 Ag Silver 107.868	48 Cd Cadmium 112.411	49 In Indium 114.818	50 Sn Tin 118.71	51 Sb Antimony 121.760	52 Te Tellurium 127.6	53 I Iodine 126.904	54 Xe Xenon 131.29
55 Cs Cesium 132.905	56 Ba Barium 137.327	57-71 Lanthanide Series	72 Hf Hafnium 178.49	73 Ta Tantalum 180.948	74 W Tungsten 183.85	75 Re Rhenium 186.207	76 Os Osmium 190.23	77 Ir Iridium 192.22	78 Pt Platinum 195.08	79 Au Gold 196.967	80 Hg Mercury 200.59	81 Tl Thallium 204.383	82 Pb Lead 207.2	83 Bi Bismuth 208.980	84 Po Polonium 209	85 At Astatine 209	86 Rn Radon 222
87 Fr Francium 223	88 Ra Radium 226	89-103 Actinide Series	104 Rf Rutherfordium [261]	105 Db Dubnium [262]	106 Sg Seaborgium [266]	107 Bh Bohrium [264]	108 Hs Hassium [269]	109 Mt Meitnerium [268]	110 Ds Darmstadtium [281]	111 Rg Roentgenium [272]	112 Cn Copernicium [285]	113 Uut Ununtrium [288]	114 Fl Flerovium [289]	115 Uup Ununpentium [291]	116 Lv Livermorium [293]	117 Uus Ununseptium [294]	118 Uuo Ununoctium [296]
57 La Lanthanum 138.905	58 Ce Cerium 140.12	59 Pr Praseodymium 140.908	60 Nd Neodymium 144.24	61 Pm Promethium [145]	62 Sm Samarium 150.36	63 Eu Europium 151.964	64 Gd Gadolinium 157.25	65 Tb Terbium 158.925	66 Dy Dysprosium 162.50	67 Ho Holmium 164.930	68 Er Erbium 167.26	69 Tm Thulium 168.934	70 Yb Ytterbium 173.04	71 Lu Lutetium 174.967			
89 Ac Actinium 227	90 Th Thorium 232	91 Pa Protactinium 231	92 U Uranium 238	93 Np Neptunium 237	94 Pu Plutonium 244	95 Am Americium 243	96 Cm Curium 247	97 Bk Berkelium 247	98 Cf Californium 251	99 Es Einsteinium 252	100 Fm Fermium 257	101 Md Mendelevium 258	102 No Nobelium 259	103 Lr Lawrencium [262]			
		Alkali Metal	Alkaline Earth	Transition Metal	Basic Metal	Semimetal	Nonmetal	Halogen	Noble Gas	Lanthanide	Actinide						

**Experiment no (1)**  
**Separation and Analysis of First Group**  
**(Ag<sup>+1</sup>, Hg<sub>2</sub><sup>+2</sup>, Pb<sup>+2</sup>)**

Group I are consists of Silver Ag<sup>+1</sup>, Lead Pb<sup>+2</sup>, and Mercurous Hg<sub>2</sub><sup>+2</sup> and these ions are common of this group.

The chemical characteristics of the metals to be considered in this course shows that the chlorides of the three ions, Ag<sup>+1</sup>, Hg<sub>2</sub><sup>+2</sup> and Pb<sup>+2</sup> are insoluble whereas those of the other cations are soluble. It is possible, therefore, to separate these three metals from the others in a general unknown by adding Cl<sup>-</sup> to the solution to precipitate the chlorides of lead, silver, and mercurous.

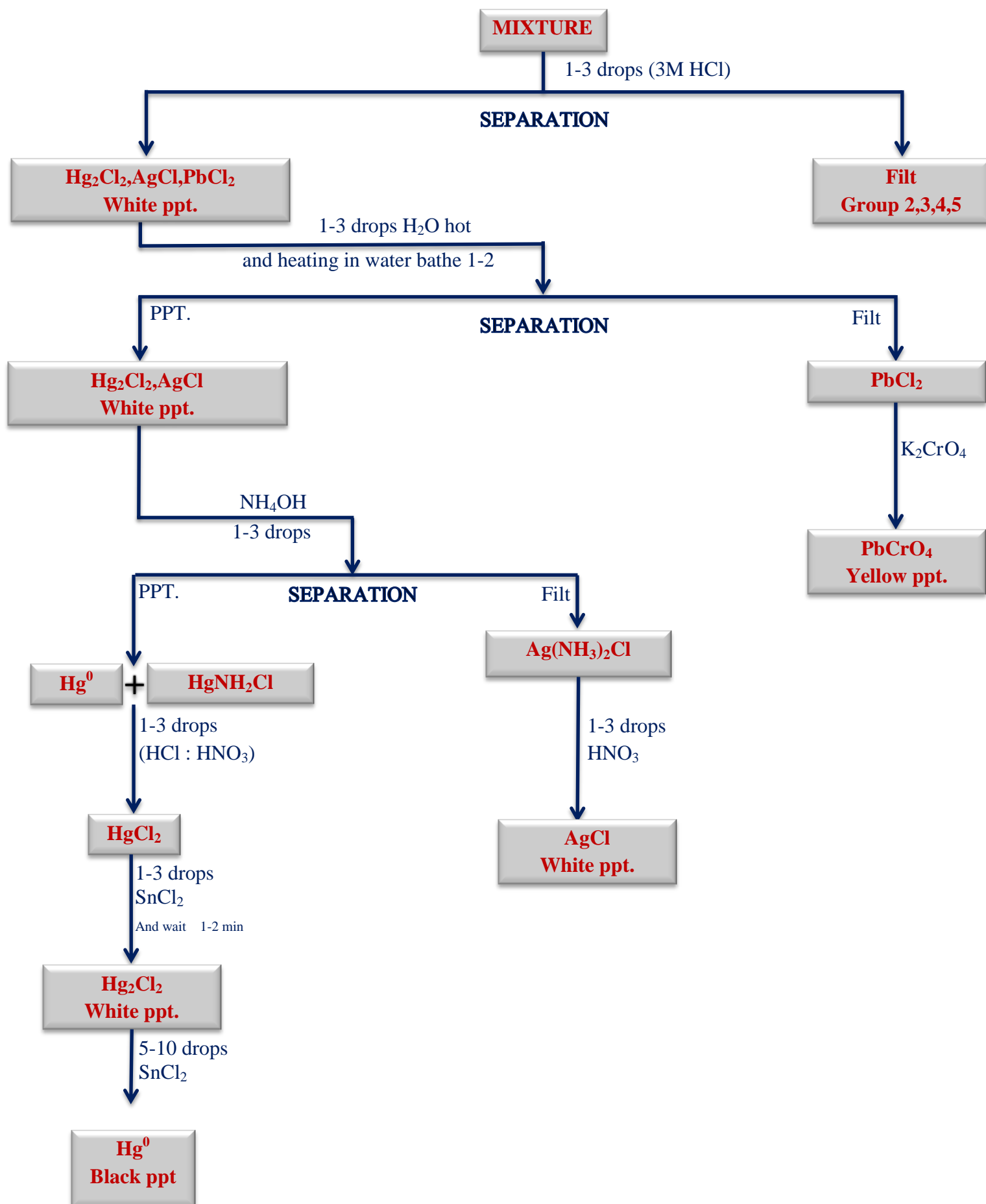
These ions precipitated by the use of an acid solution of hydrochloric acid at a concentration (3 M) these precipitations (AgCl, PbCl<sub>2</sub>, and Hg<sub>2</sub>Cl<sub>2</sub>) formed as shown in the equations below:

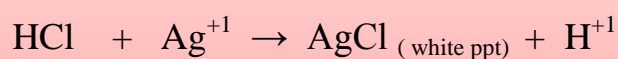
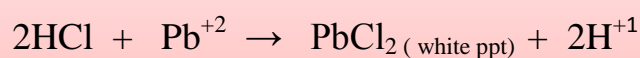
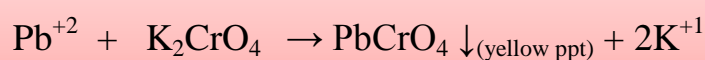
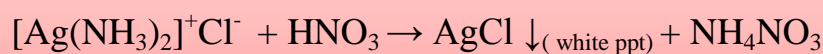
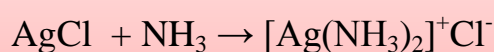


**Procedure:**

- 1- transfer 1 ml of mix. to test tube then add 3 drops of dil. HCl (3M).
- 2- stir the mix and put it in the centrifuge (2 min) then separate .
- 3- add to the filtrate 1 drop of dil. HCl .
- 4- the precipitate contain AgCl , PbCl<sub>2</sub> , Hg<sub>2</sub>Cl<sub>2</sub> which are white precipitate .
- 5- add 1 ml of hot dist. water then transfer to water bath (1-2 min.) .
- 6- transfer the test tube to centrifuge while its hot , separate the filtrated from the precipitate .
- 7- each ion will be identify by adding the specific reagent:
  - A: add K<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub> to hot filtrated while contains Pb<sup>+2</sup>, Cl<sup>-</sup> ( yellow ppt. ) .
  - B: add (NH<sub>4</sub>OH) to ppt.( AgCl , Hg<sub>2</sub>Cl<sub>2</sub>) to dissolve AgCl.
  - C: add dil. HNO<sub>3</sub> to filtrate solution to precipitate AgCl .
  - D: dissolve the ppt. of Hg<sub>2</sub>Cl<sub>2</sub> in the ( aqua regia ) then add SnCl<sub>2</sub> (white ppt. ) then change to gray after add excess of SnCl<sub>2</sub> .

## SEPARATION AND ANALYSIS OF THE FIRST GROUP I



معادلات الكشف العام لأيونات الطائفة الاولىالكشف التأكيدي لأيون الرصاصالكشف التأكيدي لأيون الفضةالكشف التأكيدي لأيون الزئبقوز