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P3

Physical_Chemistry_2nd_YUGS_EV_ST

55/100
Fifty
June

Name of a student خديجة كاظم عباس Signature [Signature] No. B6
Mustansiriyah University 2nd SEM-2026_Bologna_Process
Department of Chemistry Mid_Exam_Class_A_Paper_B

Q1/ MCQ test (Answer the following) (Marks 50 %)

1: Which two variables does the Gibbs phase rule consider as independent??
Answer: a) p & T b) F & T c) p & conc. d) T & conc.

2: If NaCl is added to ice, which property of the solution decreases?
Answer: a) LP b) VP c) FP d) BP

3: At what pressure do the three phases of CO₂ coexist in the phase diagram?
Answer: a) at 1 atm b) over 1 atm c) below 1 atm d) at any pressure

4: Which phase corresponds to a supercooled substance?
Answer: a) gas b) liquid c) solid d) plasma

5: How many phases are present when a one-component system has two degrees of freedom?
Answer: a) zero b) 1 c) 2 d) 3

6: The Clausius equation can be applied to which of the following phase equilibria?
Answer: a) melt. & freez. b) frees. & melt. c) vap. & cond. d) all of these

7: What is the relationship between the VP of a solution and the solute molality?
Answer: a) direct b) inverse c) disordered d) none of these

8: Which type of solute, when added to a solvent, alters its colligative properties?
Answer: a) non-volatile solute b) volatile solute c) pure solute d) pure solvent

9: In osmosis, the solvent moves toward which component?
Answer: a) solute b) impure solute c) mixture d) pure solvent

10- One of the most important applications of measuring molar mass of the solute is to study the change in ---.
Answer: a) m b) Π c) V d) p

Q2/ 0.5 mol of a non-P-solute was added to 12.0 mol of P-solvent, VP^* is 12.0 kPa at 295 K. What is the VP at 295 K? Determine the deviation of this solution from Raoult's law where $VP_{ideal} = 10$ kPa. (Marks 25%)

Q3/ Plot the phase diagram of the system (α and β) assumed that (α and β) do not react with each other. α freezes at (-7 °C) and β freezes at (10 °C), and that a eutectic mixture is formed when the ratio is 30 wt % of β and that the eutectic melts at (-10 °C), then label all the parts (p & F) of the diagram using the appropriate phase rule? (Marks 25%)

Q2 / 0.5 mol, 12 ml قطري , VP* 12.0 kPa , T = 295K , VP البخاري T = 295K

VP ideal = 10 kPa
الضغط البخاري المثالي

$$P_A = P_A^* \cdot T$$

$$\frac{P_A - P_A^*}{P_A^*} = \frac{n_B}{n_A}$$

$$\frac{12 - 10}{10} = \frac{0.5 \text{ mol}}{12 \times 0.1}$$

$$\frac{12 - P_A}{10} = 0.042$$

$$12 - P_A = 10 \times 0.042$$

$$12 - P_A = 0.42$$

$$P_A = 12 - 0.42 \rightarrow P_A = 11.58$$

الكل يتغير الحجم والضغط مع P ideal

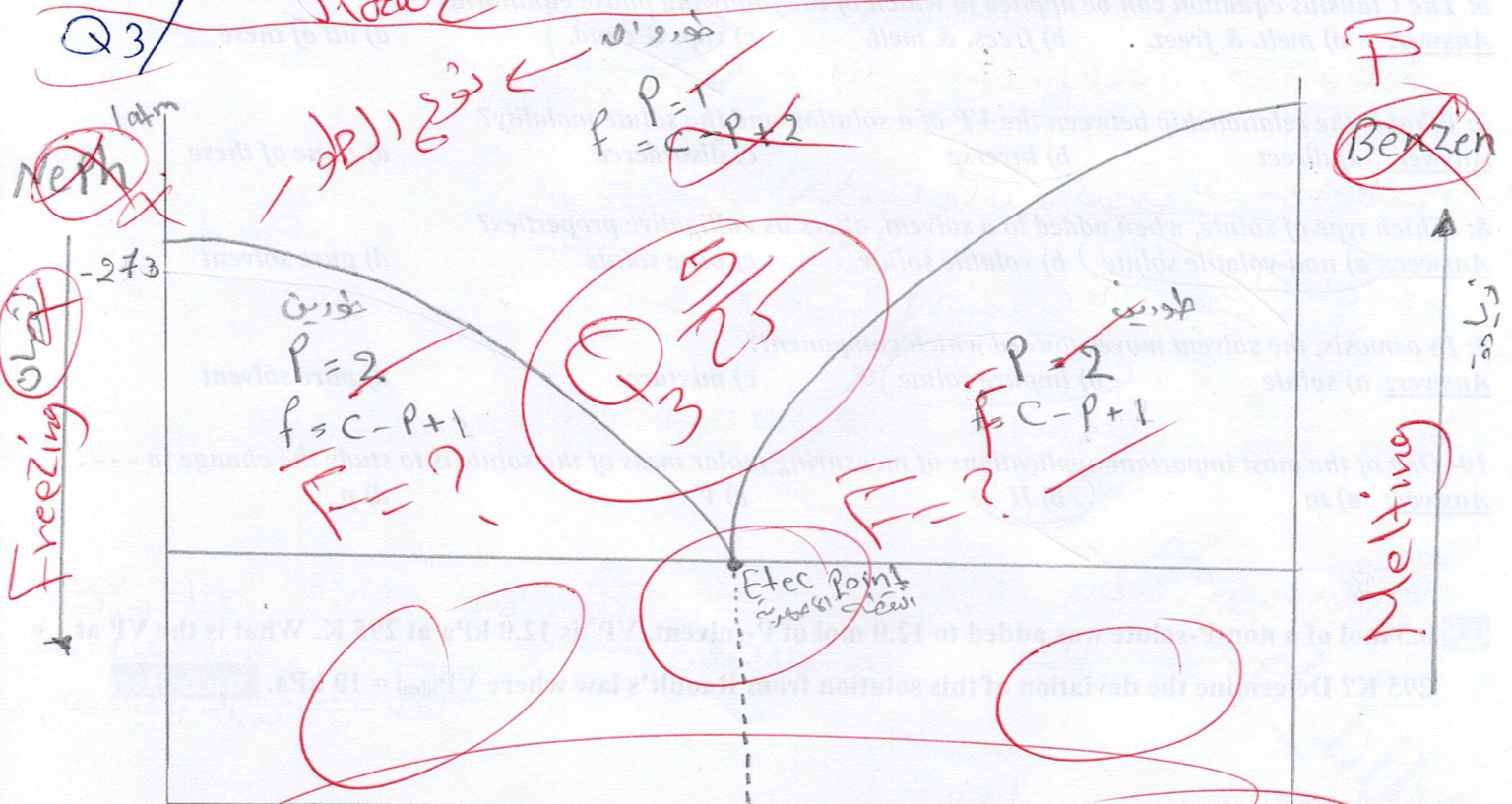
Q2 15/25

? = 0.04

P real > P ideal

هذا يعني ان الضغط الكلي هو مجموع كل الضغطين

Q3/



نوع الطور

$$F = C - P + 1$$

$$P = 2$$

$$F = C - P + 1$$

$$F = ?$$

$$P = 2$$

$$F = C - P + 1$$

$$F = ?$$

Etec Point

?