



Name of a student _____ Signature _____ No. _____

Q2/m = 4.0 g, V = 0.5 dm³, P = 103 torr, T = 344 + 273 = 307 K

$PV = nRT$

$P = \frac{103}{760} = 0.135$

$0.135 \times 0.5 \text{ dm}^3 = n \times 0.082 \times 307 \text{ K}$

$n = \frac{0.135 \times 0.5}{0.082 \times 307} \Rightarrow n = 0.0027 \text{ mol}$

$M = \frac{m}{n} \Rightarrow M = \frac{4 \text{ g}}{0.0027 \text{ mol}} \Rightarrow M = 1481 \text{ g/mol}$

Q20
2/26

C = 2, F = 4 - P

P = 1, F = 3
P = 2, F = 1

Q3

Two component system ()

